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the following solutions Stoichiometry

problems: 1. How many grams of silver

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chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100.

mL of 0.400 M potassium chromate? 2

$\text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2\text{KNO}_3(\text{aq})$

0.150 L AgNO_3

0.500 moles AgNO_3 1 moles Ag_2CrO_4

331.74 g Ag_2CrO_4

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Rb+2 left 2.50 L 3.00 L 0.45 mol Rb
0.25mol Rb total volume initial moles

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moles used $2 \times 2 + - = - + + = 0.0363 \text{ M}$

Solution Stoichiometry • An unknown diprotic acid reacts completely with 35.2 mLs of 0.45 M NaOH. How many moles of the acid were present? $\text{H}_2\text{A}(\text{aq}) + \text{NaOH}(\text{aq}) \rightarrow \text{NaH}_2\text{A}(\text{aq}) + \text{H}_2\text{O}(\text{l})$
 $\text{H}_2\text{A}(\text{aq}) + 2\text{NaOH}(\text{aq}) \rightarrow \text{Na}_2\text{A}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$

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the following solutions Stoichiometry problems: 1. How many grams of silver chromate will precipitate when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate? 2

$$\text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$$

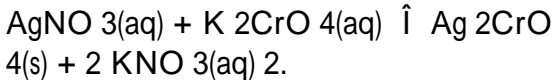
2. How many mL of 0.

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Stoichiometry Involving Solutions
Worksheet. 1. Calculate the number of mL
of 2.00 M HNO_3 solution required to react

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with 216 grams of Ag according to the equation. $3 \text{ Ag(s)} + 4 \text{ HNO}_3\text{(aq)} \text{ -----} \rightarrow$
 $3 \text{ AgNO}_3\text{(aq)} + \text{ NO(g)} + 2 \text{ H}_2\text{O(l)}$ 2.

Calculate in mL the volume of 0.500 M NaOH required to react with 3.0 grams of acetic acid.

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Worksheet

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CHEMISTRY 110 . last first . 1] How many grams of calcium phosphate can be produced from the reaction of 2.50 L of 0.250 M Calcium chloride with an excess of phosphoric acid?

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Calculate the molarity of the H_2SO_4 solution if it takes 40.0 mL of H_2SO_4 to neutralize 0.364 g of Na_2CO_3 . 0.0859 M H_2SO_4 . Back to top; Stoichiometry

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(Worksheet) Thermochemistry
(Worksheet)

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Strategy: A Write the balanced chemical equation for the reaction and calculate the number of moles of base needed to neutralize the ascorbic acid. B Using mole

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ratios, determine the amount of ascorbic acid consumed. Calculate the mass of vitamin C by multiplying the number of moles of ascorbic acid by its molar mass.

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the following solutions Stoichiometry problems: 1. How many grams of silver chromate will 7. What minimum number of grams of oxalic acid monohydrate, $\text{H}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$, would you specify for a titration of no fewer than 15.0 mL

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15 6 Answers

As we learned previously, double replacement reactions involve the reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these reactions occur in aqueous

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amounts of reactants or products in aqueous solutions. As we learned previously, double replacement reactions involve the reaction between ionic compounds in solution and, in the course of the reaction, the ions in the two reacting compounds are “switched” (they replace each other). Because these

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foot building has enough office space to accommodate 86 full-time faculty members and 167 full-time staff.

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Stoichiometry expresses the quantitative relationship between reactants and products in a chemical equation.

Stoichiometric coefficients in a balanced equation indicate molar ratios in that reaction. Stoichiometry allows us to

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predict certain values, such as the percent yield of a product or the molar mass of a gas.. Created by Sal Khan

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