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M, N and Na - Chemical Quantities ~~What are Physical Quantities?~~
Significant Figures - A Fast Review! Converting Between Grams
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Chapter 10 Chemical Quantities Study

Chemistry Chapter 10 Chemical Quantities. STUDY. PLAY. Mole.
(mol) of a substance is 6.02×10^{23} representative particles of
that substance and is the SI unit for measuring the amount of a
substance. Avogadro's Number. the number of representative
particles in a mole, 6.02×10^{23} . Representative Particle.

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Chapter 10 Chemical Quantities. mole. Avogadro's number.
representative particle. molar mass. the SI unit for measuring the
amount of a substance. number of representative particles in a mole,
 6.02×10^{23} . refers to the species present in a substance: usually
atoms, m. the mass of one mole of a pure substance.

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Quantities Avogadro's Number One important property of a mole is that it means a definite number of particles just like a dozen means a number of particles. While a dozen is only 12 particles a mole is a larger number □ 6.02×10^{23} particles.

Ch 10 - Avogadro's Number.docx - Name Date Class Chapter ...

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Chemistry Section 10 Study Guide Chemical Quantities

Chapter 10 Chemical Quantities 91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure

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matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287-289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER
(pages 287-296)

Chemical Quantities THE MOLE AND QUANTIFYING MATTER

10.1 The Mole: A Measurement of Matter Essential Understanding

The mole represents a large number of very small particles.

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another substance participating in the same reaction. For a related

section, see Equation Stoichiometry Problems with Mixtures on our

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CHAPTER 10: Chemical Quantities BASICS: □ The basic unit that is used to determine the amount of a chemical substance is called a mole □ A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance □ The mole was founded by a scientist named Avagadro, and he

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